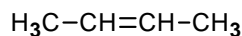
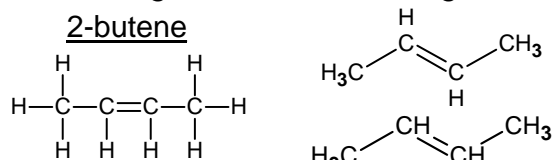


## Naming Hydrocarbons (nomenclature)

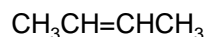


Handout: Hydrocarbons: IUPAC names

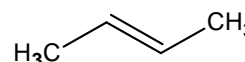
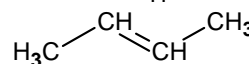
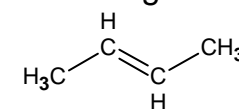
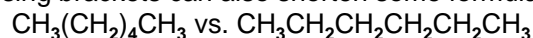
## Drawing structures: it's all good



This is called the  
"condensed structure"



Using brackets can also shorten some formulas:



On a test, choose a  
method that shows all Hs

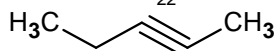
## Background: formulas for HCs

- Alkanes =  $\text{C}_n\text{H}_{2n+2}$ , enes =  $\text{C}_n\text{H}_{2n}$ , ynes =  $\text{C}_n\text{H}_{2n-2}$
- Remember enes, then think of what would happen if bond was single or triple instead.
- Provides some useful information (e.g. for compositional analysis, or to check work)
- Cannot always tell hydrocarbon type based on numbers (e.g. propyne vs. propadiene)

Q - how many hydrogens in each of these:

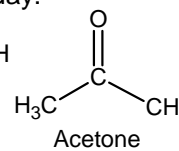
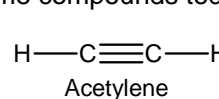
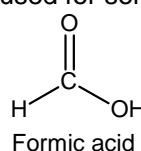
6 carbon alkane

Alkene:  $\text{C}_{22}\text{H}$



## Naming: common vs. IUPAC

- Common names used in the 1800's are still used for some compounds today:



- The International Union of Pure and Applied Chemistry (IUPAC) was established in 1900s
- Frequent revisions to nomenclature
- Systematic method allows an infinite number of compounds to be named given a few rules

## Basic names of hydrocarbons

- Hydrocarbon names are based on: 1) class 2) # of C, 3) side chain type and 4) position
  - 1) name will end in -ane, -ene, or -yne
  - 2) the number of carbons is given by a "Prefix"
- 1 meth- 2 eth- 3 prop- 4 but- 5 pent-  
6 hex- 7 hept- 8 oct- 9 non- 10 dec-

Q - What names would be given to these:

7C, 9C alkane

2C, 4C alkyne

1C, 3C alkene

## Mnemonic for first four prefixes



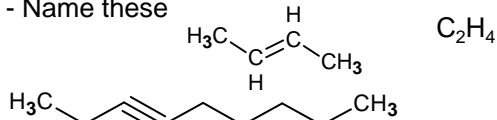
### First four prefixes

- |                 |                 |
|-----------------|-----------------|
| • <u>M</u> eth- | <u>M</u> onkeys |
| • <u>E</u> th-  | <u>E</u> at     |
| • <u>P</u> rop- | <u>P</u> eeled  |
| • <u>B</u> ut-  | <u>B</u> ananas |

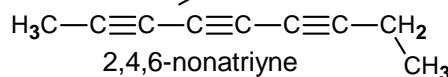
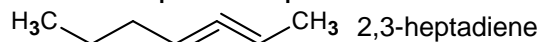
## Numbering carbons

Q- draw pentene

- Thus, naming compounds with multiple bonds is more complex than previously indicated
- Only if 2+ possibilities exist, are #s needed
- Always give double bond the lowest number
- Q - Name these



## Multiple multiple bonds



- Give 1<sup>st</sup> bond (1<sup>st</sup> point of difference) lowest #
- include di, tri, tetra, penta, etc. before ene/yne
- Comma between #s, hyphen between #-letter
- You do not need to know ene + yne

